CHEMISTRY OF MATERIALS
Additions and Corrections

Correction to Stable Cycling of Lithium Batteries Using Novel Boronium-Cation-Based Ionic Liquid Electrolytes [*Chem. Mater.* **2010,** *22,* 1038. DOI: 10.1021/cm9019815]. Thomas Rüther,* Thuy D. Huynh, Junhua Huang, Anthony F. Hollenkamp, E. Alan Salter, Andrzej Wierzbicki, Kayla Mattson, Adam Lewis, and James H. Davis, Jr.*

Page 1042. The text in the paragraph that begins "Conductivities for the boronium-based RTILs..." should read as follows:

The estimated conductivities are only slightly lower than that of the commercial (Merck) C4mpyr-NTf_2 measured under the same conditions but which has a smaller cation formula weight (142.26 g mol^{-1}). The lowest conductivity was estimated for **3b**, which was expected because this cation has the highest formula weight (261.19 g mol^{-1}) among the three materials. Among the boronium RTILs, **3c** exhibits the highest conductivity (1.56 mS cm^{-1}) owing to a relatively low formular weight (145.07 g mol^{-1}) and structural features like lower symmetry and flexibility, giving rise to a greater degree of freedom.[15,16,23c,26a] With a greater degree of charge delocalization than in conventional counterparts,[39] **3a** displays only a slightly lower conductivity despite its considerably high formular weight (177.03 g mol^{-1}) when compared to **3c** and C4mpyr-NTf_2. These factors may account for the conductivities when comparing the boronium based RTILs with C4mpyr-NTf_2.

Page 1043. The unit on the Y-axis of Figure 5 must be I/Acm^{-2}.

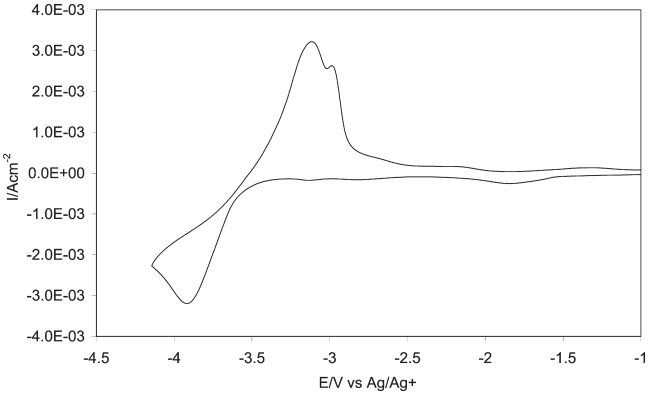

Figure 5

Page 1044. Part b of Figure 6 is missing and should be as appears here.

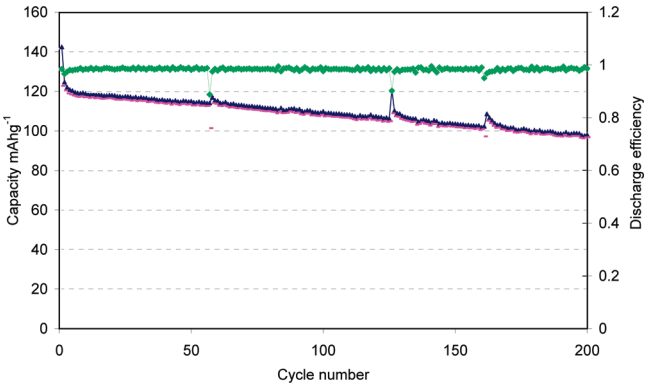

Figure 6. Part b.

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